

Name:

Period:

# H1 Stability in Bonding

Complete each of the following items. Use your textbook on pages 602-621

**Directions:** Complete the table below by using the formula of each compound to identify the elements that each compound contains and the number of atoms of each of these elements in a unit of the compound. The first formula has been done for you. 3 pts

Formula	Element 1	Element 2	Element 3
H <sub>2</sub> O	2 HYDROGEN	1 OXYGEN	----
a. NaOH			
b. NaCl			
c. NH <sub>4</sub>			
d. H <sub>2</sub> SO <sub>4</sub>			
e. SiO <sub>2</sub>			

**Directions:** Review each of the chemical formulas below. Write the name of the chemical using the rules discussed and used in class. Refer to your notes and textbook for help! 2 pts

a. 9. H<sub>2</sub>O; \_\_\_\_\_

b. 10. HCl; \_\_\_\_\_

c. 11. AgCl; \_\_\_\_\_

d. 12. CaSO<sub>4</sub>; \_\_\_\_\_

e. 13. FeSO<sub>4</sub>; \_\_\_\_\_

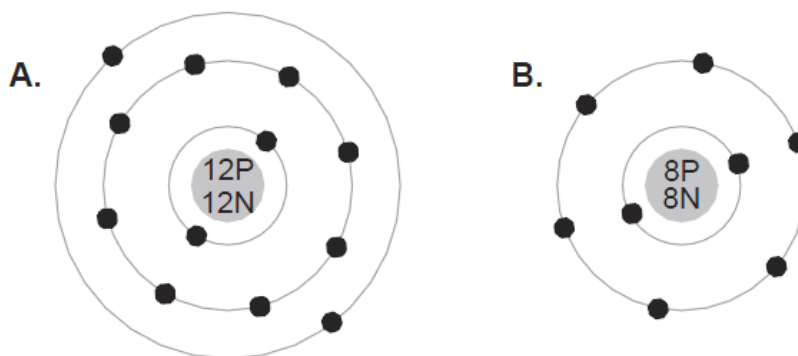
Name:

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Directions: *Write the chemical formulas for each of chemical names listed below:*

- a. sodium chloride \_\_\_\_\_
- b. manganese sulfide \_\_\_\_\_
- c. potassium oxide \_\_\_\_\_
- d. copper (iii) bromide \_\_\_\_\_
- e. lithium carbonate \_\_\_\_\_
- f. carbon tetrachloride \_\_\_\_\_
- g. calcium sulfide \_\_\_\_\_
- h. nitrogen trihydride \_\_\_\_\_
- i. potassium permanganate \_\_\_\_\_

**Directions:** *Study the diagram below. Write your answers to the questions in the spaces provided. 3 pts*



What element is atom A? \_\_\_\_\_ B? \_\_\_\_\_

**If atom A loses electrons to atom B,**

- a. how many electrons will atom A lose? \_\_\_\_\_
- b. how many electrons will atom B gain? \_\_\_\_\_
- c. what will be the oxidation number of atom A? \_\_\_\_\_
- d. what will be the oxidation number of atom B? \_\_\_\_\_
- e. what will be the total charge of the compound formed? \_\_\_\_\_
- f. what type of bond will form? \_\_\_\_\_